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and lipids (all of which will be defined later in this chapter) that are the fundamental molecular components of all organisms.

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2.1 The Nature of Matter What 3 subatomic particles make up atoms?

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## **Life**

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Chapter 2 Vocabulary. Atom Nucleus

Electron Element Isotope Compound

Ionic bond Ion Covalent bond Molecule.

Van der Waals forces Hydrogen bond

Cohesion Adhesion Mixture Solution

Solute Solvent Suspension pH scale. Acid

Base Buffer Monomer Polymer

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Carbohydrate Monosaccharide Lipid  
Nucleic acid Nucleotide.

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2.3 Carbon Compounds Lesson

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Objectives Describe the unique qualities of carbon. Describe the structures and functions of each of the four groups of macromolecules. Lesson Summary The Chemistry of Carbon Organic chemistry is the study of compounds with bonds between carbon atoms. Carbon atoms have four valence electrons, allowing them to form

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PLAY. A solution with a pH value of 12 is a \_\_\_\_\_. Base. Atoms are composed of \_\_\_\_\_. Protons, Neutrons, and Electrons. The attraction between a water molecule and another water molecule is called \_\_\_\_\_. Cohesion.

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A B; atom: the basic unit of matter:  
nucleus: the center of the atom:  
electron: a negatively charged ...

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Biology: Chapter 2, The Chemistry of  
Life. Description. Molecules, Matter,  
Atoms, Ionic and Covalent Bonds. Total

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Cards. 32. Subject. Biology. Level. 9th Grade. Created. ... (this means that H<sub>2</sub>O is water, but H<sub>3</sub>O is not) Compounds have different physical and chemical properties from the elements they are made of (For example, Chlorine (Cl ...

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what to pull off subsequently mostly

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## **Crossword Puzzle Answer Key**

CHAPTER OUTLINE 2.1: The Building Blocks of Molecules 2.2: Water 2.3: Biological Molecules Foods such as bread, fruit, and cheese are rich sources of biological macromolecules. KEY BUILDING BLOCKS OF LIVING

ORGANISMS Key elements : •Carbon (C)  
•Hydrogen (H) •Nitrogen (N) •Oxygen

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(O) • Sulfur (S) • Phosphorus (P) Form :  
• Carbohydrates • Nucleic acids • Proteins  
• Lipids

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